

The Crystal Structure of Palladium *n*-Propyl Mercaptide

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The crystal structure of palladium-*n*-propyl mercaptide has been determined from Patterson and Fourier syntheses using three-dimensional counter data and refined by least squares. The final *R* index is 0.053 2833 observed reflexions. The space group is $P\bar{1}$ and the cell constants are (at 75 °F) $a = 13.766 \pm 0.004$ Å, $b = 10.564 \pm 0.004$ Å, $c = 11.499 \pm 0.004$ Å, $\alpha = 98.36 \pm 0.02^\circ$, $\beta = 111.40 \pm 0.02^\circ$, $\gamma = 102.83 \pm 0.02^\circ$. The molecule exists as a hexamer $\{\text{Pd}(\text{SC}_3\text{H}_7)_2\}_6$ in the crystal. The palladium atoms form a six-membered puckered ring. Each adjacent pair of metal atoms is joined by a double mercaptide bridge which is folded along the sulphur-sulphur axis. Each palladium is surrounded by four sulphur atoms in an approximately square planar configuration.

Introduction

This work forms part of a programme of study of the structures of various metal mercaptides. Results of the structure analysis of mercury mercaptides by Bradley & Kunchur (1964, 1965) and Kunchur (1964) show that the sulphur of the mercapto group acts as a bridge between two mercury atoms. This gives rise to interesting structural arrangements in these compounds. Mercaptides of palladium were prepared by Mann & Purdie (1935) and Hayter & Humiec (1964). The ethyl and phenyl derivatives were found to be highly insoluble and hence linear polymeric structures for these compounds were proposed. The compounds of palladium with *n*-propyl and higher mercaptans are, however, soluble in organic solvents and can be readily crystallized. Preliminary work on single crystals of palladium *n*-propylmercaptide showed that the unit cell is triclinic and contains six formula units of this compound (Mann & Purdie, 1935). This is consistent with the results of the molecular weight determination of this compound in solution, which showed the molecular complexity *n* to have a value approaching six. On the basis of these data, Hayter & Humiec (1964) proposed two possible structures for palladium *n*-propyl mercaptide, both of which are based on square planar palladium (II).

In one structure palladium atoms are at the corners of a regular octahedron and the bridging mercaptide groups lie above each edge. In the other possibility, the palladium atoms form a six-membered planar ring, each adjacent pair of metals being joined by a double mercaptide bridge, which is folded along the sulphur-sulphur axis. A detailed structure analysis of this compound was undertaken to discover the correct one of these possibilities, and to define the molecular geometry.

Experimental

The crystals of palladium *n*-propyl mercaptide, kindly supplied by Dr Hayter, were small, well defined orange-red needles along the *b*-axis. The unit-cell dimensions and space group of this compound were determined

by Mann & Purdie (1935). Their values are: $a = 14.3$, $b = 10.3$, $c = 11.3$ Å, $\alpha = 98^\circ$, $\beta = 113.5^\circ$, $\gamma = 101^\circ$. There are six formula units of this compound in the unit cell. The space group could be $P1$ or $P\bar{1}$, but, the piezoelectric test being negative, they assigned the space group $P\bar{1}$ to the compound.

In the present work, the unit-cell dimensions and the space group were redetermined on the G.E. XRD-6 counter diffractometer using Cu $K\alpha$ radiation. The results are as follows: palladium *n*-propyl mercaptide $\{\text{Pd}(\text{S.C}_3\text{H}_7)_2\}_6$. Molecular weight: 1539.5. Triclinic: $a = 13.766 \pm 0.004$ Å, $b = 10.564 \pm 0.004$ Å, $c = 11.499 \pm 0.004$ Å, $\alpha = 98.36 \pm 0.02^\circ$, $\beta = 111.40 \pm 0.02^\circ$, $\gamma = 102.83 \pm 0.02^\circ$. The dimensions were obtained using high order axial reflexions under the 'fine conditions' of the diffractometer, at 75 °F. The volume of the unit cell is 1470.2 Å³. The absorption coefficient for X-rays, $\lambda = 1.5404$ Å, $\mu = 188.0$ cm⁻¹. The density = 1.73 g.cm⁻³, calculated on the basis of six formula units of the monomeric compound in the unit cell, agrees approximately with the experimental value of 1.75 g.cm⁻³. Total number of electrons in the unit cell $F(000) = 768$.

The space group $P\bar{1}$ was assumed to be correct on the basis of the work of the previous authors. The crystals employed for the unit-cell measurements and data collection were needles of dimensions 0.06 × 0.06 × 0.05 mm. For the linear absorption coefficient $\mu = 188.0$ cm⁻¹, the value for μr is 0.56. ρ -absorption corrections were applied. Data were collected on the G.E. XRD-6 diffractometer with scintillation counter using Cu $K\alpha$ radiation and a balanced pair of cobalt and nickel filters. The stationary-crystal stationary-counter technique was used. All 2929 possible independent reflexions within a range $0^\circ < 2\theta < 100^\circ$ were examined. Of these, 96 were too weak to be measured. α -splitting and the usual 1/*Lp* corrections were applied and the raw intensities were converted into F^2 and F .

Structure determination

A three-dimensional 'unsharpened' Patterson synthesis was computed. In the space group $P\bar{1}$, there are eight centres of symmetry and since there are six palladium

atoms in the unit cell, the job of fixing the palladium positions was rather complicated. The structure was solved by use of Buerger's 'image seeking' method. A palladium-palladium single weight peak was located and a line was drawn from the origin to the center of the peak. This line was used as an image to locate the remaining palladium atoms. No attempt was made to locate the sulphur atoms at this stage.

Three cycles of structure-factor least squares were carried out for the six palladium atoms. This gave an agreement index $R=0.31$. A palladium-phased three-dimensional electron density Fourier series was then computed. This yielded the positions of all the twelve sulphur and two carbon atoms. The agreement index R at this stage was 0.21. A second electron density difference Fourier series was computed, using the coefficients obtained by subtracting the contributions of all palladium and sulphur atoms from the F_{obs} . This gave the coordinates of all but two carbon atoms. The agreement index R for all the atoms located up to this stage was 0.16. Another electron density difference Fourier series was computed; this gave the remaining two carbon atoms. The structure, now essentially determined, was refined by a series of least-squares analyses of the three-dimensional data. Four cycles, with isotropic atomic temperature factors, were carried out on the IBM 7040 computer, using the program written by Okaya and modified by Kartha & Harris. The agreement index R for all observed reflexions fell to 0.11. Another four cycles of refinement with individual anisotropic temperature factors and a block-diagonal approximation, reduced the R to 0.053. A $1/f$ weighting with f corresponding to the contribution of sulphur

was adopted in the final stages of the least-squares analysis. This weighting is approximately equivalent to differential synthesis weighting. Atomic scattering factors for palladium, sulphur, and carbon were taken from the *International Tables for Crystallography*, Vol. III (1962). The anomalous corrections $\Delta f'$ and $\Delta f''$ for palladium with Cu $K\alpha$ were applied. The atomic shifts in the final cycle were much smaller than the standard deviations of the atomic positions, and so the refinement was considered to be complete. An electron density difference Fourier series, computed using the final coordinates of all the atoms except the hydrogens, was free from details of any significance. No attempt was made to locate the hydrogen atoms; the calculations described above do not take into account the contributions of hydrogen atoms to the structure factors. The final agreement index R was 0.053 for all observed reflexions and 0.058 if all unobserved reflexions were included. The final atomic and thermal parameters as well as the F_{obs} and F_{calc} are those of the final least-squares cycle. The atomic and thermal parameters are listed in Tables 1 and 7 while the F_{obs} and F_{calc} are listed in Table 2.

Discussion and description of the structure

The various bond lengths, bond angles and their estimated standard deviations are shown in Table 3. The standard deviations in the atomic parameters were obtained by inverting the matrix in the final least-squares cycle. The labelling scheme used to describe the palladium *n*-propyl mercaptide molecule can be seen in Fig. 1. For every atom in the unit cell, there is another atom

Table 1. Fractional final coordinates and their standard deviations (\AA)

	x/a	y/b	z/c	σ_x	σ_y	σ_z
Pd(1)	0.16796	0.01108	0.27509	0.00070	0.00138	0.00089
Pd(2)	0.01487	0.25220	0.88932	0.00072	0.00140	0.00090
Pd(3)	0.18416	0.26365	0.16557	0.00070	0.00138	0.00088
S(1)	0.0583	0.3800	0.0945	0.0026	0.0049	0.0032
S(2)	0.3049	0.1390	0.2302	0.0024	0.0049	0.0032
S(3)	0.1538	0.2218	0.3455	0.0028	0.0051	0.0033
S(4)	0.8245	0.1997	0.8081	0.0026	0.0051	0.0032
S(5)	0.9739	0.1179	0.6877	0.0026	0.0052	0.0032
S(6)	0.2018	0.2844	0.9750	0.0026	0.0054	0.0033
C(1)	0.948	0.327	0.1469	0.013	0.0251	0.0163
C(2)	0.967	0.422	0.270	0.013	0.025	0.015
C(3)	0.882	0.377	0.317	0.020	0.039	0.025
C(4)	0.307	0.038	0.087	0.010	0.023	0.013
C(5)	0.420	0.111	0.079	0.016	0.033	0.019
C(6)	0.517	0.099	0.182	0.016	0.043	0.022
C(7)	0.285	0.308	0.480	0.012	0.023	0.014
C(8)	0.308	0.457	0.509	0.021	0.029	0.023
C(9)	0.581	0.477	0.373	0.022	0.038	0.023
C(10)	0.782	0.269	0.663	0.011	0.023	0.014
C(11)	0.668	0.190	0.562	0.012	0.026	0.015
C(12)	0.582	0.198	0.608	0.016	0.040	0.028
C(13)	0.080	0.034	0.702	0.012	0.022	0.015
C(14)	0.162	0.105	0.663	0.010	0.025	0.013
C(15)	0.245	0.031	0.671	0.013	0.036	0.020
C(16)	0.259	0.467	0.996	0.012	0.028	0.017
C(17)	0.578	0.341	0.937	0.020	0.040	0.036
C(18)	0.624	0.484	0.924	0.019	0.036	0.030

Table 2. Observed and calculated structure factors
The unobserved reflections are marked by an asterisk.

Table with 20 columns (H, K, L, FOBS, FCAL) and multiple rows of data. Asterisks mark unobserved reflections.

Table 2 (cont.)

Table with 20 columns: H, K, L, FOBS, FCAL (repeated 5 times). Each column contains numerical data, possibly representing coordinates or parameters. The table contains a dense grid of numbers with some asterisks and negative signs interspersed.

Table 2 (cont.)

Table with 20 columns (H, K, L, FOBS, FCAL) and multiple rows of numerical data. The data is organized in groups, with asterisks (*) marking specific rows. The values range from -10 to 12 across various columns.

related to it by a centre of inversion. Thus, for convenience of identification, Pd(1) for example, is related to Pd(1') by a centre of symmetry. A diagrammatic representation of the structure projected down the c^* and b^* axes is shown in Figs. 3 and 4 respectively. From these diagrams it becomes obvious that in this crystal structure the compound exists not as a monomer but as a hexamer consisting of six Pd(S.C₃H₇)₂ fragments. The basic unit of this large molecule consists of one palladium and four sulphur atoms arranged in an approximately square planar configuration. Six units of this type are linked to each other by the sharing of two sulphur atoms between every two adjacent palladium atoms. The sulphur atoms thus act as a bridge between two palladium atoms. The Pd-S-Pd bridge angles vary between 79.7 and 85.8°. The nature of the sharing of sulphur atoms can be seen in Fig. 2(a). In the resulting structure then, there are two six-membered sulphur rings which are S(1), S(3), S(5'), S(6'), S(2'), S(4) and S(1'), S(3'), S(5), S(6), S(2), S(4'). These are packed approximately above each other and one

ring is related to the other by a centre of inversion. The average distance between the two rings is 3.09 Å. There is also a ring of six palladium atoms which is

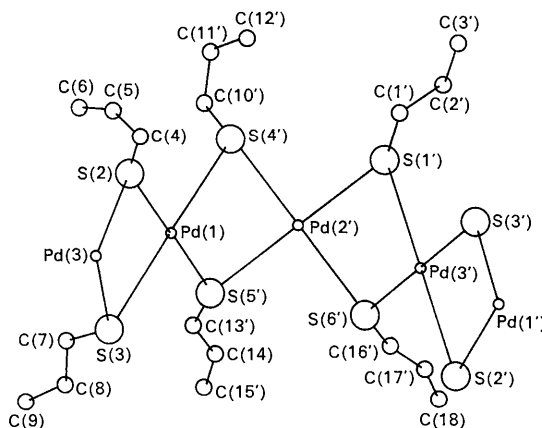


Fig. 1. Schematic diagram of a part of the molecule of palladium *n*-propyl mercaptide.

Table 3. *Dimensions of the molecule*

	Bond length (Å)	σ (Å)		Angles (°)	σ (°)
Pd(1)-S(2)	2.305	0.004	S(5)-Pd(1)-S(4)	82.0	0.2
Pd(1)-S(3)	2.399	0.004	S(5)-Pd(1)-S(3)	98.2	0.2
Pd(1)-S(4)	2.413	0.004	S(5)-Pd(1)-S(2)	177.8	0.2
Pd(1)-S(5)	2.314	0.005	S(4)-Pd(1)-S(3)	176.6	0.2
			S(4)-Pd(1)-S(2)	97.6	0.2
Pd(2)-S(1)	2.450	0.005	S(3)-Pd(1)-S(2)	82.1	0.2
Pd(2)-S(4)	2.358	0.004	S(6)-Pd(2)-S(5)	98.7	0.2
Pd(2)-S(5)	2.454	0.004	S(6)-Pd(2)-S(4)	174.6	0.2
Pd(2)-S(6)	2.331	0.005	S(6)-Pd(2)-S(1)	80.3	0.2
			S(5)-Pd(2)-S(4)	80.3	0.2
Pd(3)-S(1)	2.282	0.004	S(5)-Pd(2)-S(1)	177.9	0.2
Pd(3)-S(2)	2.288	0.004	S(4)-Pd(2)-S(1)	100.5	0.2
Pd(3)-S(3)	2.303	0.005	S(6)-Pd(3)-S(3)	174.0	0.2
Pd(3)-S(6)	2.308	0.005	S(6)-Pd(3)-S(2)	94.7	0.2
			S(6)-Pd(3)-S(1)	84.5	0.2
S(1)-C(1)	1.82	0.02	S(3)-Pd(3)-S(2)	84.6	0.2
C(1)-C(2)	1.60	0.03	S(3)-Pd(3)-S(1)	96.0	0.2
C(2)-C(3)	1.46	0.04	S(2)-Pd(3)-S(1)	177.4	0.2
S(2)-C(4)	1.92	0.02	C(1)-S(1)-Pd(3)	113.3	0.7
C(4)-C(5)	1.62	0.03	C(1)-S(1)-Pd(2)	110	0.7
C(5)-C(6)	1.46	0.04	C(2)-C(1)-S(1)	110	1.4
S(3)-C(7)	1.89	0.05	C(3)-C(2)-C(1)	110	2.0
C(7)-C(8)	1.52	0.03	C(4)-S(2)-Pd(1)	104.9	0.6
C(8)-C(9)	1.59	0.04	C(4)-S(2)-Pd(3)	110.5	0.6
S(4)-C(10)	1.81	0.02	C(5)-C(4)-S(2)	104	1.4
C(10)-C(11)	1.58	0.03	C(4)-C(5)-C(6)	115	2.2
C(11)-C(12)	1.48	0.04	C(7)-S(3)-Pd(3)	105	1.6
S(5)-C(13)	1.84	0.02	C(7)-S(3)-Pd(1)	108	0.6
C(13)-C(14)	1.43	0.03	C(8)-C(7)-S(3)	116	1.7
C(14)-C(15)	1.50	0.03	C(7)-C(8)-C(9)	113	2.3
S(6)-C(16)	1.88	0.02	C(10)-S(4)-Pd(1)	99	0.6
C(16)-C(17)	1.47	0.04	C(10)-S(4)-Pd(2)	107	0.6
C(17)-C(18)	1.54	0.04	C(11)-C(10)-S(4)	114	1.4
			C(10)-C(11)-C(12)	110	201
			C(13)-S(5)-Pd(1)	112	0.6
			C(13)-S(5)-Pd(2)	110	0.7
			C(14)-C(13)-S(5)	113	1.4
			C(13)-C(14)-C(15)	111	2.0
			C(16)-S(6)-Pd(3)	100	0.7
			C(16)-S(6)-Pd(2)	105	0.7
			C(17)-C(16)-S(6)	115	1.8
			C(16)-C(17)-C(18)	112	2.8

sandwiched between the two sulphur rings in an anti-prismatic configuration. Mean planes passing through the palladium and sulphur atoms were calculated to see if these rings were planar. From the results of these calculations it can be seen that (see Table 6) the rings are not planar but puckered. The dimensions of the rings are shown in Figs. 2(a) and 2(b). The sulphur-sulphur bonds in the sulphur ring vary between 3.38–3.70 Å and the S–S–S angles between 114.1°–124.2°. In the palladium ring the palladium distances vary between 3.014 and 3.235 Å and the angles between 116.3° and 123.7°.

The n-propyl groups to which sulphur is bonded are located outside the framework of palladium and sulphur atoms. Since the carbon-sulphur bond is a single bond, the n-propyl groups can rotate around this bond

and orient themselves to sterically suitable positions in the crystal structure. The geometry of the molecule of (Pd n-propyl mercaptide)₆ is very interesting. The molecule has a cage structure and has a large hole of 7 Å diameter within it.

The Pd–S distances vary between 2.282 and 2.454 Å. Mean planes were calculated to test the planarity of the PdS₄ units. The deviations of a number of atoms from the mean planes are significant in all the three independent PdS₄ units. Of the six different carbon-sulphur bonds, three are normal, the lengths being about 1.82 Å. The remaining three bonds have somewhat larger value (see Table 3). The agreement between the carbon-carbon distances in the six independent n-propyl groups is not very satisfactory. But considering that the carbon atoms were located in the presence of heavy palladium and sulphur atoms, little significance need be attached to these deviations.

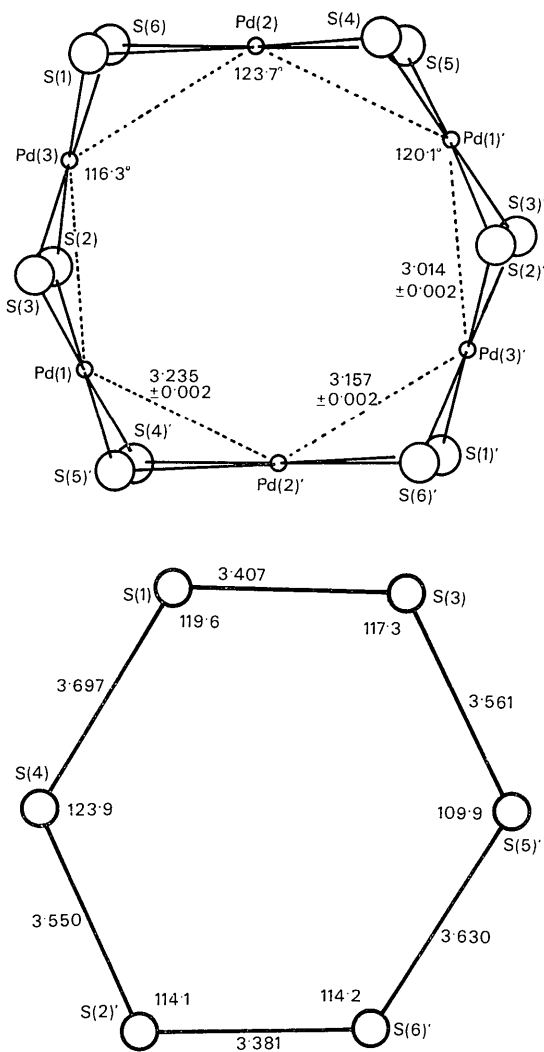


Fig. 2. A diagram showing (a) the dimensions of the ring of six palladium atoms and the sharing of the ring of six sulphur atoms between them. (b) The dimensions of the ring of six sulphur atoms.

Table 4. Intermolecular approaches

	Length (Å)	σ (Å)
C(6')—C(6)	4.39	0.04
C(12')—C(18)	4.02	0.04
C(9) —C(18)	5.10	0.04
C(9) —C(9')	4.39	0.04
C(9) —C(18')	4.27	0.04
C(6) —C(12)	4.66	0.04
C(6) —C(18')	4.10	0.04
C(9) —C(15)	5.33	0.04
C(9) —C(12)	4.55	0.04
C(7) —C(15)	3.81	0.04
C(16) —C(18)	4.53	0.04
C(6) —C(9)	4.19	0.04

The manner in which the various groups are packed in the crystal structure can be seen in Figs. 3 and 4. The various intermolecular close approaches are listed in Table 4. The short intermolecular approaches are those between the terminal methyl groups of the n-propyl chains, and their values vary between 3.81 and 5.33 Å. None of their values are smaller than the van der Waals diameter of the methyl group.

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Table 5. Sulphur bridge angles

Pd(1)–S(2)–Pd(3)	82.0° ± 0.1°
Pd(1)–S(3)–Pd(3)	79.7 ± 0.1
Pd(3)–S(1)–Pd(2)	83.6 ± 0.1
Pd(3)–S(6)–Pd(2)	85.8 ± 0.1
Pd(2)–S(4)–Pd(1')	85.4 ± 0.1
Pd(2)–S(5)–Pd(1')	85.4 ± 0.1

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Table 6. *Least-squares planes and the out-of-plane distances of the plane determining atoms*

Plane number 1. This plane is defined by 5 atoms. The equation of the plane is $0.1524P + 0.9550Q - 0.2546R = 0.242$ Å. The standard deviation of the distances from the plane is 0.613 Å. Out-of-plane distances of the plane-determining atoms are:

Pd(1)	-1.130 Å
S(3)	0.495
S(4')	0.086
S(5')	-0.041
S(2)	0.590

Plane number 2. This plane is defined by 5 atoms. The equation of the plane is $0.1168P + 0.9494Q + 0.2916R = 3.738$ Å. The standard deviation of the distances from this plane is 0.306 Å. Out-of-plane distances of the plane-determining atoms are:

Pd(2)	-0.384 Å
S(1)	0.067
S(5)	-0.330
S(4)	0.334
S(6)	0.313

Plane number 3. This plane is defined by 5 atoms. The equation of the plane is $0.5168P + 0.8045Q + 0.2929R = 2.971$ Å. The standard deviation of the distances from the plane is 0.2197 Å. Out-of-plane distances of the plane-determining atoms are:

Pd(3)	0.141 Å
S(1)	0.096
S(2)	0.095
S(3)	-0.438
S(6)	0.107

Plane number 4. This plane is defined by 6 atoms. The equation of the plane is $-0.4961P + 0.8180Q - 0.2912R = 0.000$ Å. The standard deviation of the distances from the plane is 1.160 Å. Out-of-plane distances of the plane-determining atoms are:

Pd(1)	-1.683
Pd(2)	0.242
Pd(3)	0.884
Pd(1')	1.683
Pd(2')	-0.242
Pd(3')	-0.884

Plane number 5. This plane is defined by 6 atoms. The equation of the plane is $0.1147P + 0.9698Q - 0.2154R = -0.771$ Å. The standard deviation of the distances from the plane is 1.342 Å. Out-of-plane distances of the plane-determining atoms are:

S(2)	1.605 Å
S(5)	0.471
S(6)	-0.377
S(1')	-2.604
S(3')	-0.129
S(4')	1.035

Angles between the planes

Plane A	Plane B	Angle
1	3	140.6°
2	3	155.4
1	4	148.2

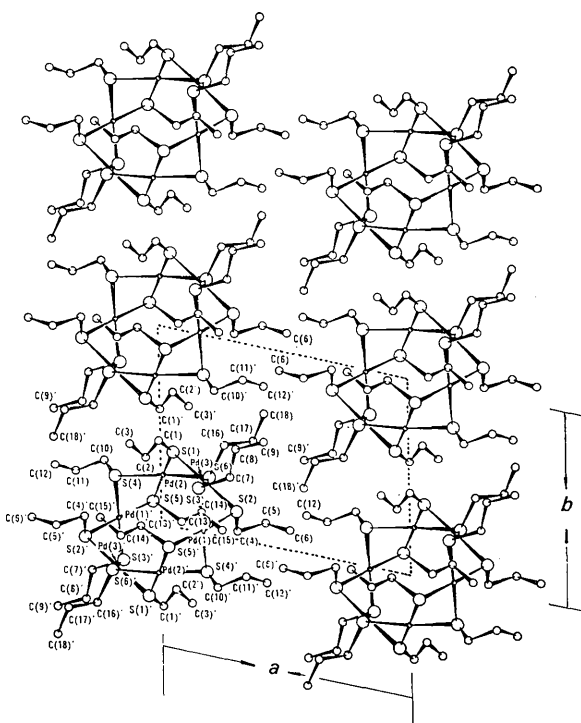


Fig. 3. A diagrammatic representation of the structure projected down the c^* axis.

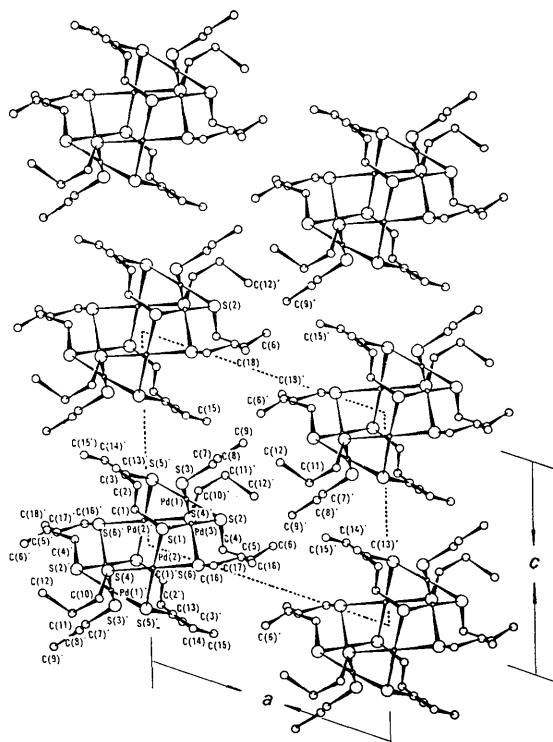


Fig. 4. A diagrammatic representation of the structure projected down the b^* axis.

Table 7. Anisotropic thermal vibration parameters

The B_{ij} are coefficients in the temperature factor expression:

$$\exp [-(B_{11}h^2 + B_{22}k^2 + B_{33}l^2 + B_{12}hk + B_{13}hl + B_{23}kl)]$$

$$U_{11} = B_{11}/2\pi^2a^{*2}; U_{12} = B_{12}/4\pi^2a^*b^*; \text{ etc.}$$

	U_{11}	U_{22}	U_{33}	U_{12}	U_{13}	U_{23}
Pd(1)	0.0399	0.0489	0.0496	0.000	-0.0015	-0.006
Pd(2)	0.0451	0.0486	0.0503	0.0000	-0.0047	0.0019
Pd(3)	0.0402	0.0489	0.0481	0.0003	-0.0026	-0.0009
S(1)	0.0494	0.0474	0.0520	0.0064	-0.0062	-0.0031
S(2)	0.0348	0.0516	0.0523	0.0005	0.0010	0.0001
S(3)	0.0457	0.0506	0.0563	0.0047	0.0109	0.0026
S(4)	0.0466	0.0551	0.0494	0.0029	-0.0078	-0.0035
S(5)	0.0444	0.0560	0.0473	-0.0020	-0.0003	0.0079
S(6)	0.0459	0.0599	0.0503	-0.0019	0.0000	0.0046
C(1)	0.0502	0.0841	0.0785	0.0046	0.0174	0.0017
C(2)	0.0683	0.0857	0.0651	0.0120	0.0176	-0.0072
C(3)	0.1049	0.1377	0.1272	0.0057	0.0290	0.0041
C(4)	0.0467	0.0890	0.0536	0.0019	-0.0019	-0.0130
C(5)	0.0879	0.1300	0.0895	-0.0247	0.0091	-0.0163
C(6)	0.0638	0.1927	0.1152	0.0340	-0.0090	0.0099
C(7)	0.0776	0.0691	0.0627	-0.0059	0.0150	-0.0005
C(8)	0.2111	0.0819	0.1163	0.0313	-0.0545	-0.0428
C(9)	0.2451	0.1225	0.0999	0.0203	-0.0841	-0.0454
C(10)	0.0747	0.0621	0.0629	-0.0063	-0.0234	0.0154
C(11)	0.0787	0.0839	0.0648	0.0032	-0.0334	-0.0021
C(12)	0.00469	0.1519	0.1819	0.0334	-0.0334	-0.0142
C(13)	0.0516	0.0604	0.0842	0.0172	0.0172	0.0082
C(14)	0.0388	0.0991	0.0570	-0.0072	-0.0015	0.0136
C(15)	0.0777	0.1553	0.1021	0.0184	-0.0138	0.0190
C(16)	0.0709	0.0874	0.0998	-0.0600	-0.0099	0.0113
C(17)	0.0840	0.0939	0.2147	-0.0096	0.0059	0.0131
C(18)	0.1264	0.1119	0.2526	-0.0426	-0.0343	-0.0337

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The Crystal and Molecular Structure of 4,4'-Diamino-3,3'-dichlorobiphenyl

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4,4'-Diamino-3,3'-dichlorobiphenyl, $C_{12}H_{10}N_2Cl_2$, crystallizes with cell dimensions $a = 12.5$, $b = 3.85$, $c = 23.7$ Å, $\beta = 108^\circ$, space group $P2_1/c$ and $Z = 4$. The structure has been determined from three-dimensional X-ray data and refined by the minimum residual method, with isotropic temperature factors for individual atoms, to give a final R index of 14%. The two phenyl rings are not coplanar, but are twisted around the linkage between them, C(1)–C(1'), so as to be mutually inclined at an angle of 21° ; in addition, each phenyl ring is bent through a small angle (approximately 2.2° and 3.4° respectively) away from the line C(1)–C(1'). The length of the bond C(1)–C(1') is 1.515 ± 0.024 Å.

Introduction and experimental

A preliminary examination of one projection of the structure of 4,4'-diamino-3,3'-dichlorobiphenyl by

Toussaint (1948) gave an electron density map with spurious symmetry in addition to the true symmetry of the actual structure. The map suggested that the molecules are probably planar, or very nearly so, with chlorine atoms in the *trans* positions. We have now completed the analysis of the structure, using three-dimensional X-ray data kindly supplied by Dr Toussaint, and find that the phenyl rings are mutually inclined at an angle of 21° .

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